

Formel	Name	Dissoziation	Säurerest	Salze	Herstellung
H <sub>2</sub> SO <sub>3</sub>	Schwefelige	H <sub>2</sub> SO <sub>3</sub> → 2H+SO <sub>3</sub>	Sulfit	Sulfite	SO <sub>2</sub> +H <sub>2</sub> O → H <sub>2</sub> SO <sub>3</sub>
H <sub>2</sub> SO <sub>4</sub>	Schwefel	H <sub>2</sub> SO <sub>4</sub> → 2H+SO <sub>3</sub>	Sulfat	Sulfate	SO <sub>3</sub> +H <sub>2</sub> O → H <sub>2</sub> SO <sub>4</sub>
H <sub>2</sub> CO <sub>3</sub>	Kohlensäure	H <sub>2</sub> CO <sub>3</sub> → 2H+CO <sub>3</sub>	Carbonat	Carbonate	CO <sub>2</sub> +H <sub>2</sub> O → H <sub>2</sub> CO <sub>3</sub>
H <sub>3</sub> PO <sub>3</sub>	Phosphorige	H <sub>3</sub> PO <sub>3</sub> → 3H+PO <sub>3</sub>	Phosphit	Phosphite	P <sub>2</sub> O <sub>3</sub> +3 H <sub>2</sub> O → H <sub>3</sub> PO <sub>3</sub>
H <sub>3</sub> PO <sub>4</sub>	Phosphor	H <sub>3</sub> PO <sub>4</sub> → 3H+PO <sub>3</sub>	Phosphat	Phosphate	P <sub>2</sub> O <sub>5</sub> +3H <sub>2</sub> O → 2H <sub>3</sub> PO <sub>4</sub>
HNO <sub>2</sub>	Salpetrige	HNO <sub>2</sub> → H+NO <sub>2</sub>	Nitrit	Nitrite	N <sub>2</sub> O <sub>3</sub> +H <sub>2</sub> O → 2HNO <sub>2</sub>
HNO <sub>3</sub>	Salpeter	HNO <sub>3</sub> → H+NO <sub>3</sub>	Nitrat	Nitrate	N <sub>2</sub> O <sub>5</sub> +H <sub>2</sub> O → 2HNO <sub>3</sub>
HCl	Chlorwasser	HCl → H+Cl	Chloride	Chloride	H <sub>2</sub> +Cl → 2HCl
HBr	Bromwasser	HBr → H+Br	Bromid	Bromide	S+H <sub>2</sub> → H <sub>2</sub> S
HI	Iodwasser	HI → H+I	Jodid	Jodide	Br+H <sub>2</sub> → 2HBr
H <sub>2</sub> S	Schwefelwa	H <sub>2</sub> S → 2H+S	Sulfid	Sulfide	I+H <sub>2</sub> → 2HI